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Write correct formulas for the products in these double replacement reactions.

1)  $\text{H}_3\text{PO}_4 + 2) \text{K}_2\text{CO}_3 + \text{BaCl}_2 \rightarrow$  3)  $\text{Al} + \text{HNO}_3 \rightarrow$  4)  $\text{Co} + \text{AgNO}_3 \rightarrow$  5)  $\text{AgNO}_3 + \text{KCl} \rightarrow$  6)  $\text{H}_2\text{SO}_4 + \text{BaCl}_2 \rightarrow$  7)  $\text{H}_2\text{SO}_4 + \text{Ca(OH)}_2 \rightarrow$  8)  $\text{H}_2\text{SO}_4 + \text{CaCO}_3 \rightarrow$  9)  $\text{H}_2\text{SO}_4 + \text{CaS} \rightarrow$  10)  $\text{Ag}_2\text{CO}_3 + \text{K}_2\text{CrO}_4 \rightarrow$

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reactions. A chemical reaction is a



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process that changes one set of chemicals into another set of chemicals. The elements or compounds that enter into the reaction are the reactants. The elements or compounds produced by the reaction are the products. Chemical reactions involve changes in the chemical bonds that join atoms in compounds. Energy in Reactions Some

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chemical reactions release energy;  
others absorb energy.

## **2.4 Chemical Reactions and Enzymes**

Chapter 11 Chemical Reactions 113  
SECTION 11.1 DESCRIBING CHEMICAL  
REACTIONS (pages 321–329) This  
section explains how to write equations

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describing chemical reactions using appropriate symbols. It also describes how to write balanced chemical equations when given the names or formulas of the reactants and products in a chemical reaction.

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in chemical reaction, all of the atoms present at the start of the reaction are present at the end of the reaction. ...

Answer:  $2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$ . What are the steps to balancing a chemical equation? 1. Write the equation 2. Count

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the atoms 3. Use coefficients to balance the atoms 4. Look back and check to make sure it is balanced correctly.

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